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ERRATUM

Foam Separation of Anions: Stoichiometry

ROBERT B. GRIEVES and DIBAKAR BHATTACHARYYA

[article in *Separation Science*, 7(2), 115-129 (1972)]

The statement at the top of p. 128 concerning the variation of the stoichiometry of the $S_2O_3^{2-}$ system is in error. If an ion exchange mechanism were occurring, S would *increase* (and not decrease) during the course of an experiment as the ratio of Br^- to $S_2O_3^{2-}$ built up. Evidently, a soluble complex was formed; however, in contrast to I^- the free surfactant and reacted surfactant were removed at different rates (Fig. 2), which could have produced shifts in the equilibrium during the course of a foaming run. Analogously to I^- , S did increase as a function of X_i/Z_i , brought about by an increase in the ratio of free to reacted surfactant, which could be predicted from the reaction $2EDTA^+ + S_2O_3^{2-} \rightleftharpoons (EHDA)_2 S_2O_3$ as X_i/Z_i were increased.