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## Separation Science and Technology

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### Erratum

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## ERRATUM

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### Foam Separation of Anions: Stoichiometry

ROBERT B. GRIEVES and DIBAKAR BHATTACHARYYA

[article in *Separation Science*, 7(2), 115-129 (1972)]

The statement at the top of p. 128 concerning the variation of the stoichiometry of the  $S_2O_3^{2-}$  system is in error. If an ion exchange mechanism were occurring,  $S$  would *increase* (and not decrease) during the course of an experiment as the ratio of  $Br^-$  to  $S_2O_3^{2-}$  built up. Evidently, a soluble complex was formed; however, in contrast to  $I^-$  the free surfactant and reacted surfactant were removed at different rates (Fig. 2), which could have produced shifts in the equilibrium during the course of a foaming run. Analogously to  $I^-$ ,  $S$  did increase as a function of  $X_i/Z_i$ , brought about by an increase in the ratio of free to reacted surfactant, which could be predicted from the reaction  $2EDTA^+ + S_2O_3^{2-} \rightleftharpoons (EHDA)_2S_2O_3$  as  $X_i/Z_i$  were increased.